

Figure 20.1
Biochemistry: A Short Course, Second Edition
 © 2013 W. H. Freeman and Company

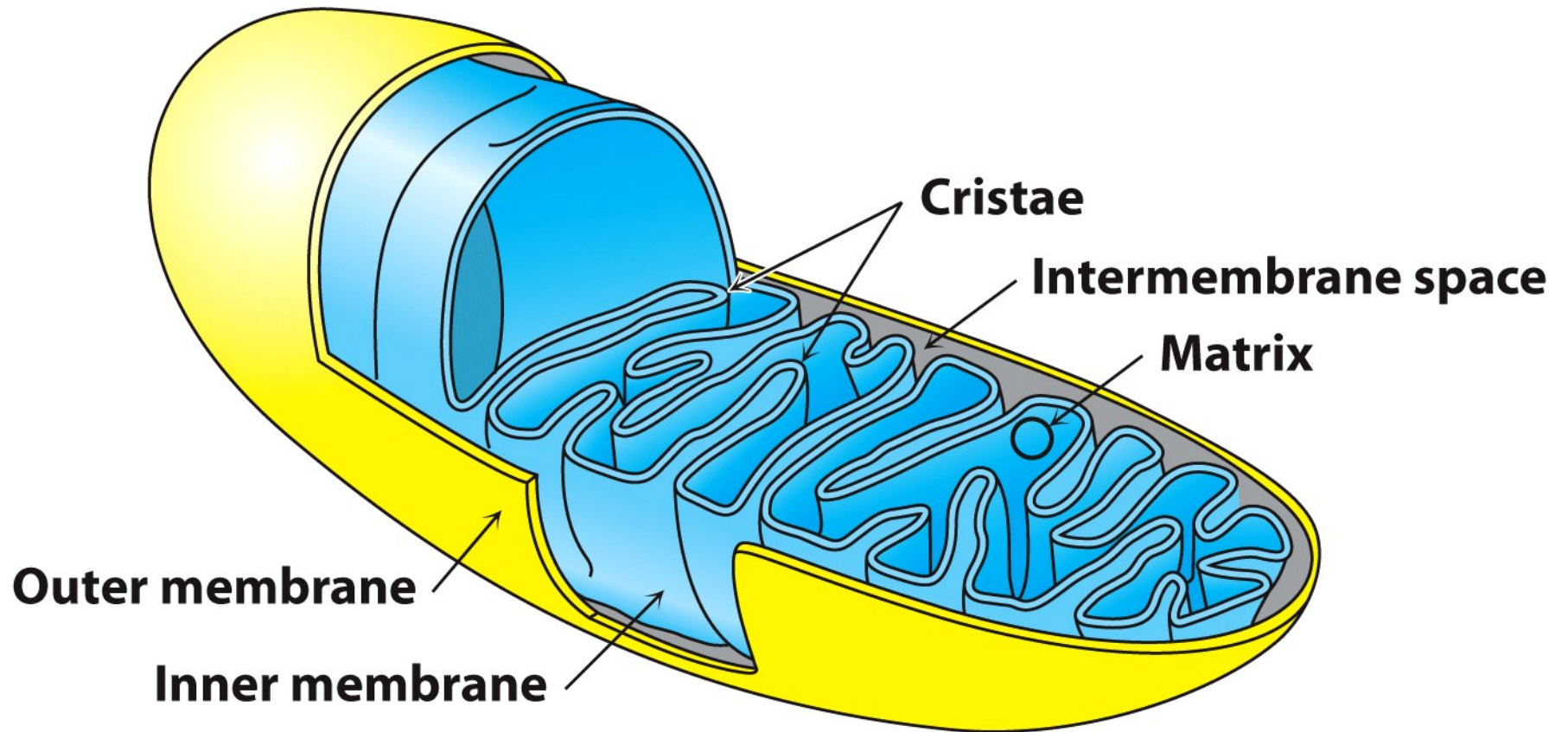
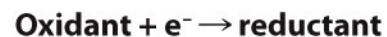


Figure 20.2b
Biochemistry: A Short Course, Second Edition
© 2013 W. H. Freeman and Company

Table 20.1 Standard reduction potentials of some reactions

Oxidant	Reductant	<i>n</i>	<i>E'</i> _o (V)
Succinate + CO ₂	α-Ketoglutarate	2	-0.67
Acetate	Acetaldehyde	2	-0.60
Ferredoxin (oxidized)	Ferredoxin (reduced)	1	-0.43
2 H ⁺	H ₂	2	-0.42
NAD ⁺	NADH + H ⁺	2	-0.32
NADP ⁺	NADPH + H ⁺	2	-0.32
Lipoate (oxidized)	Lipoate (reduced)	2	-0.29
Glutathione (oxidized)	Glutathione (reduced)	2	-0.23
FAD	FADH ₂	2	-0.22
Acetaldehyde	Ethanol	2	-0.20
Pyruvate	Lactate	2	-0.19
2 H ⁺	H ₂	2	0.00 ¹
Cytochrome <i>b</i> (+3)	Cytochrome <i>b</i> (+2)	1	+0.07
Dehydroascorbate	Ascorbate	2	+0.08
Ubiquinone (oxidized)	Ubiquinone (reduced)	2	+0.10
Cytochrome <i>c</i> (+3)	Cytochrome <i>c</i> (+2)	1	+0.22
Fe (+3)	Fe (+2)	1	+0.77
$\frac{1}{2}$ O ₂ + 2 H ⁺	H ₂ O	2	+0.82

Note: *E'*_o is the standard oxidation–reduction potential (pH 7, 25°C, except where noted) and *n* is the number of electrons transferred. *E'*_o refers to the partial reaction written as



¹Standard oxidation–reduction potential at pH = 0.

Table 20.1

Biochemistry: A Short Course, Second Edition
© 2013 W. H. Freeman and Company

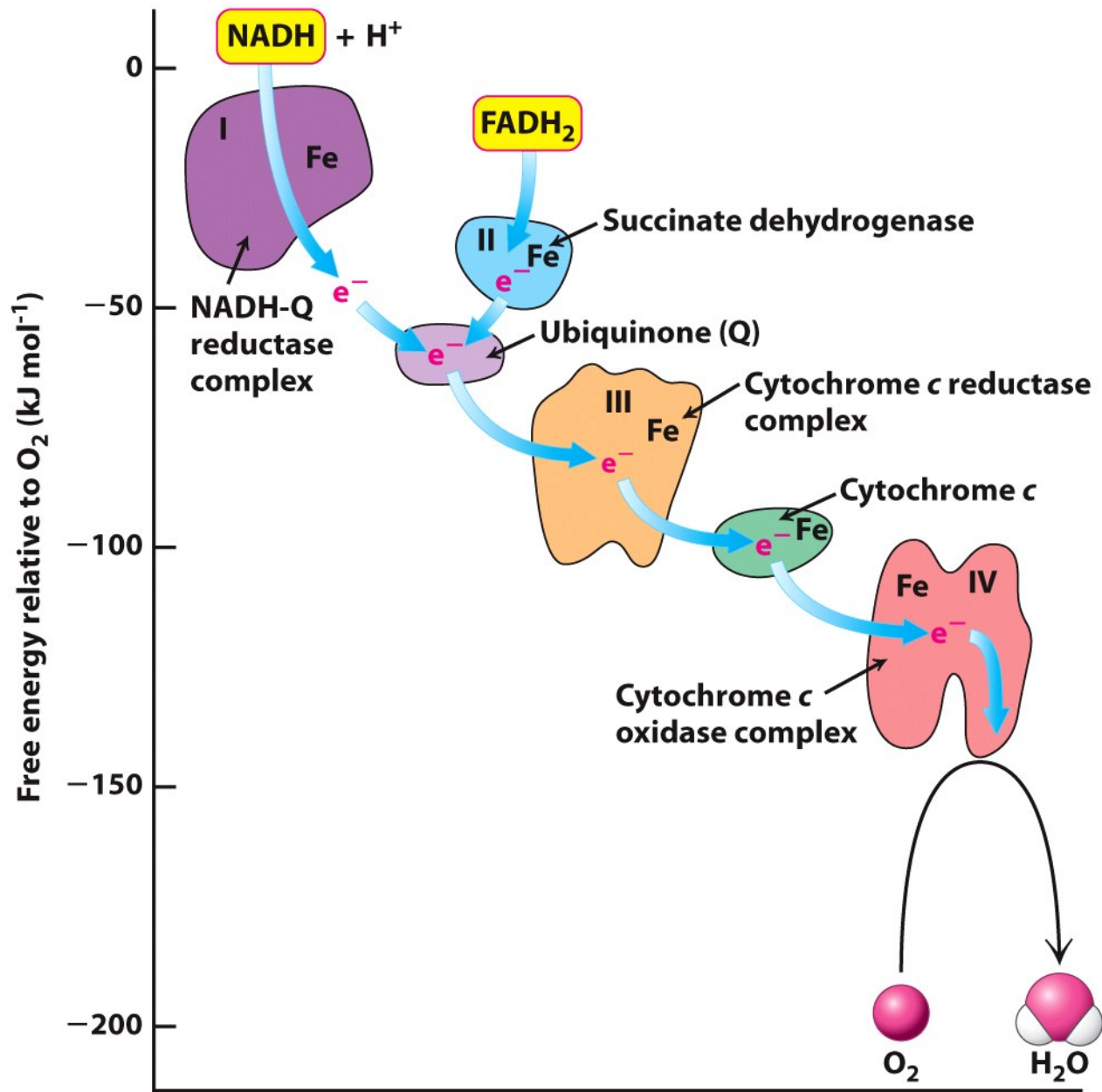


Figure 20.6
Biochemistry: A Short Course, Second Edition
 © 2013 W. H. Freeman and Company

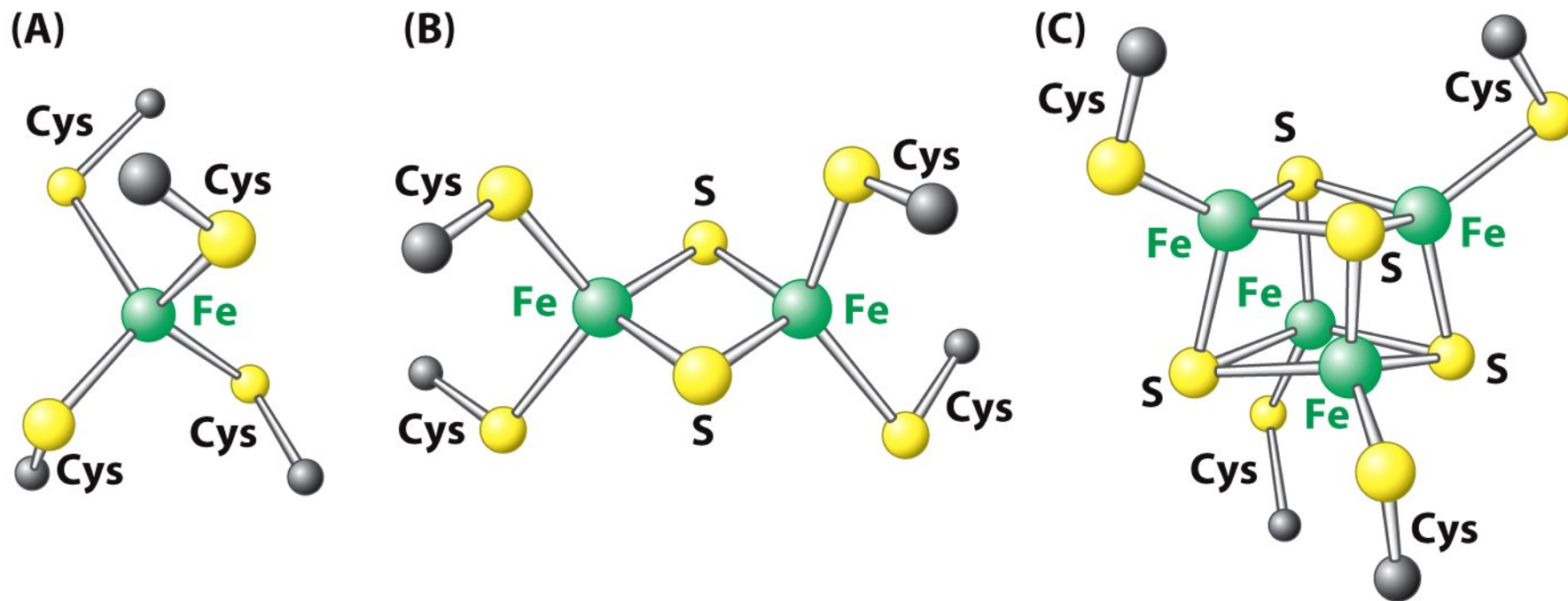


Figure 20.7
Biochemistry: A Short Course, Second Edition
© 2013 W. H. Freeman and Company

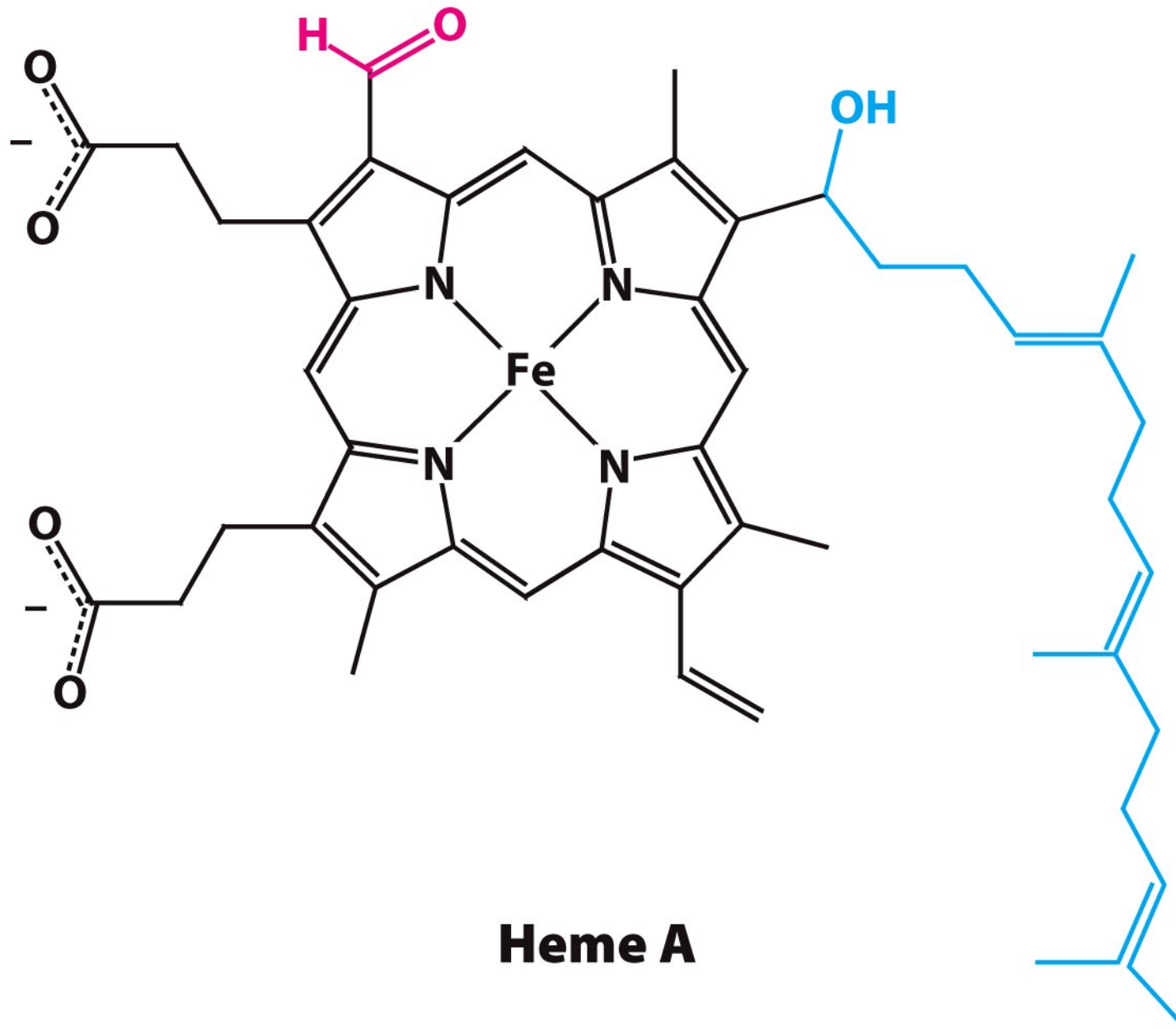
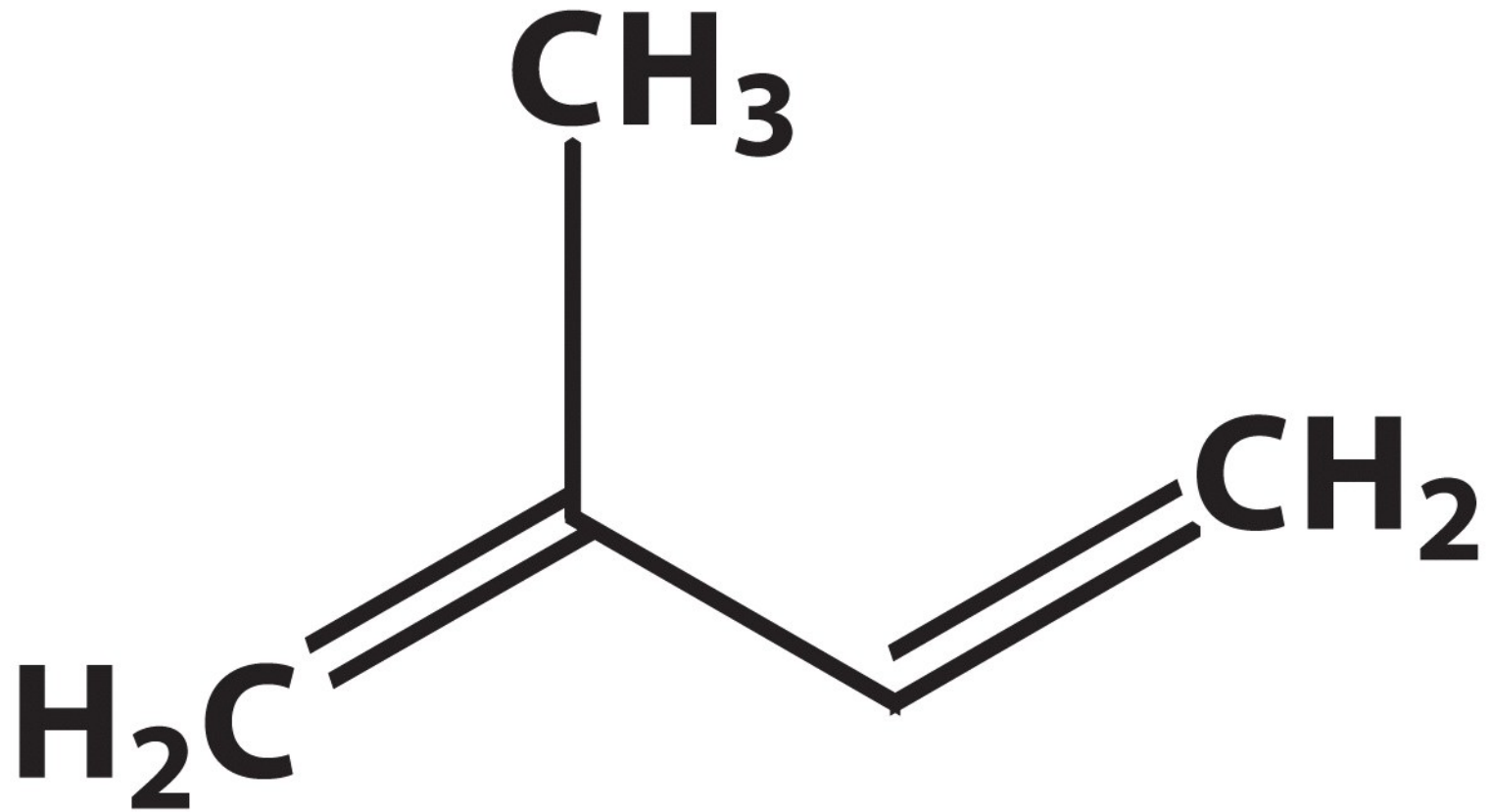


Figure 20.8
Biochemistry: A Short Course, Second Edition
© 2013 W. H. Freeman and Company



Isoprene

Unnumbered 20 p356

Biochemistry: A Short Course, Second Edition

© 2013 W. H. Freeman and Company

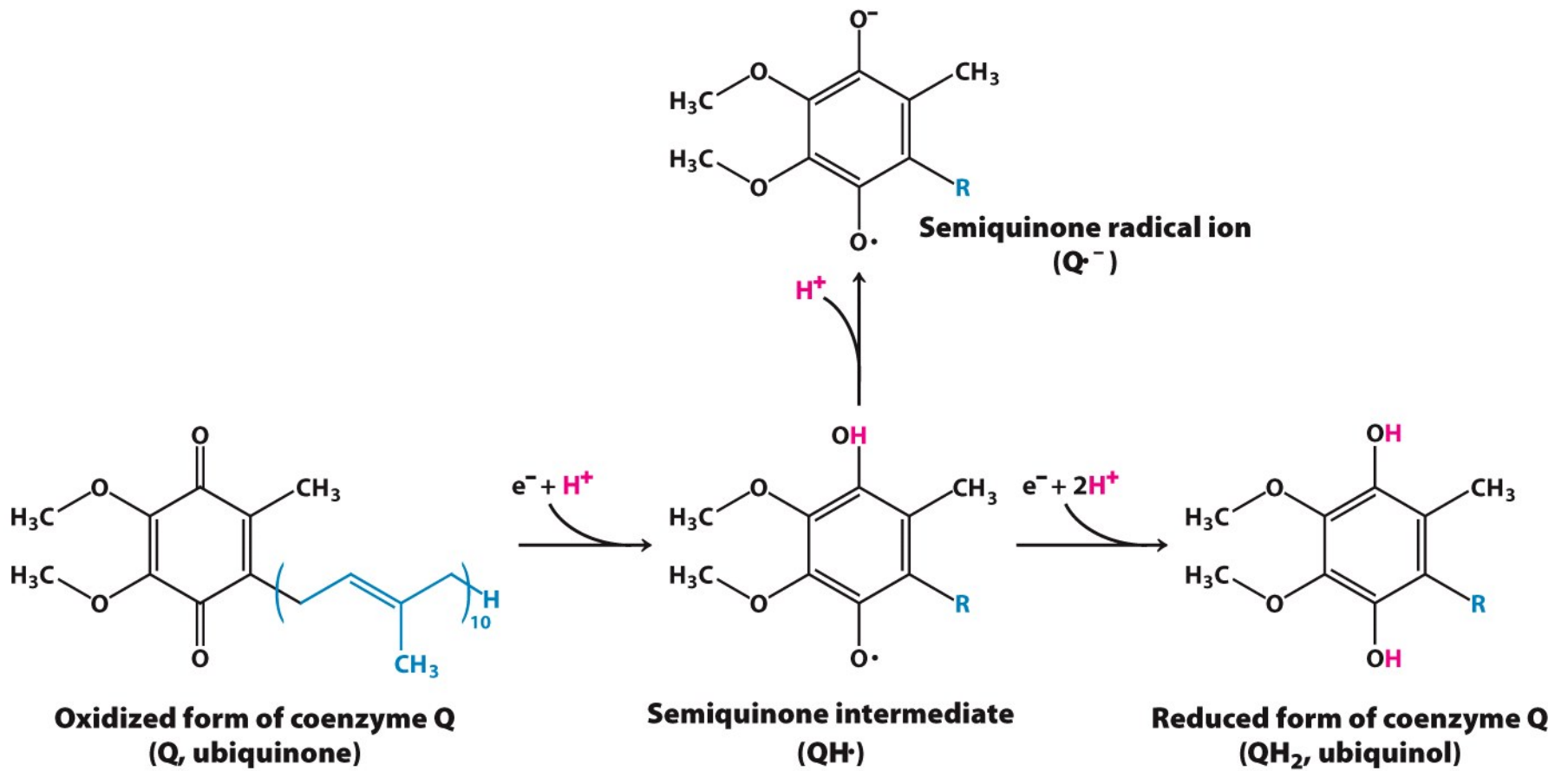


Figure 20.9
Biochemistry: A Short Course, Second Edition
 © 2013 W. H. Freeman and Company

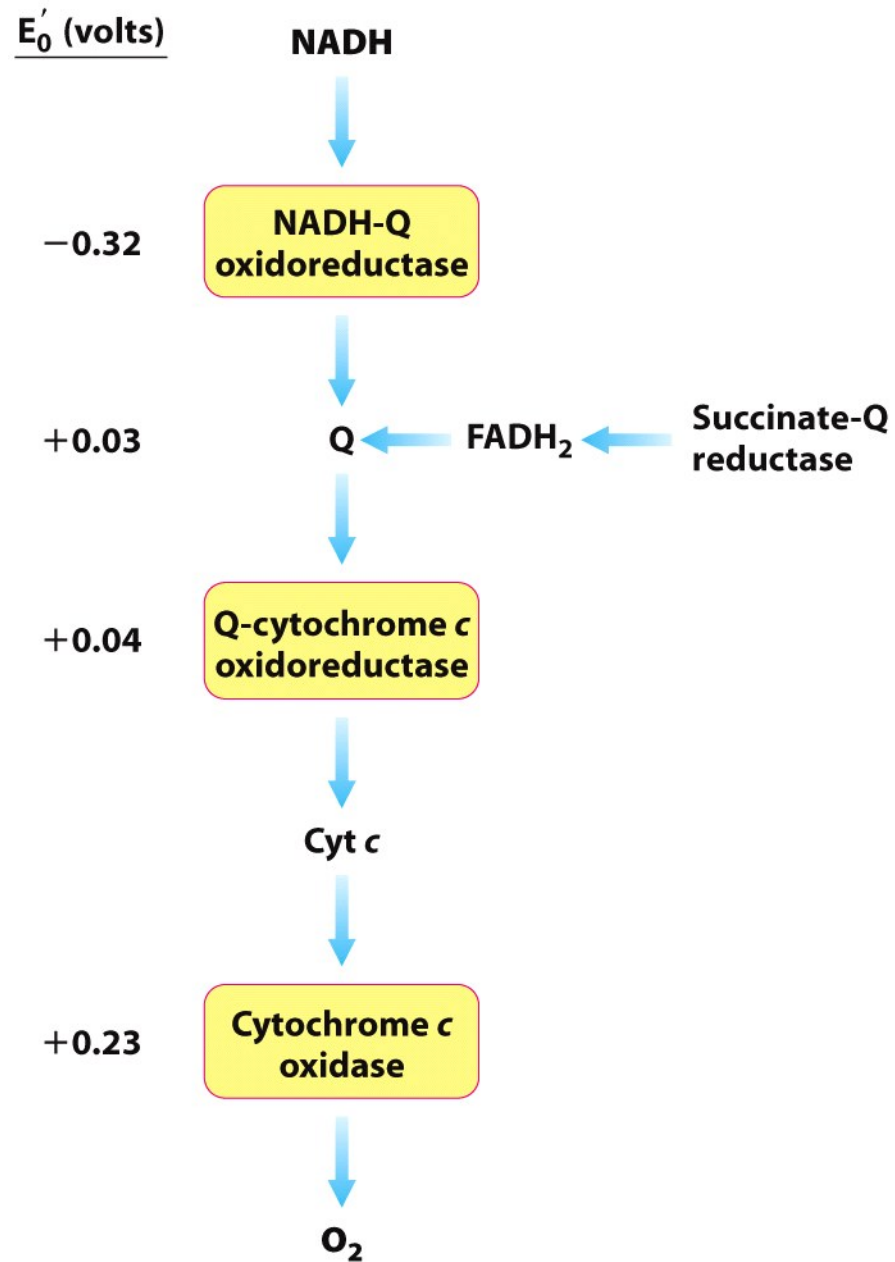


Figure 20.10
Biochemistry: A Short Course, Second Edition
 © 2013 W. H. Freeman and Company

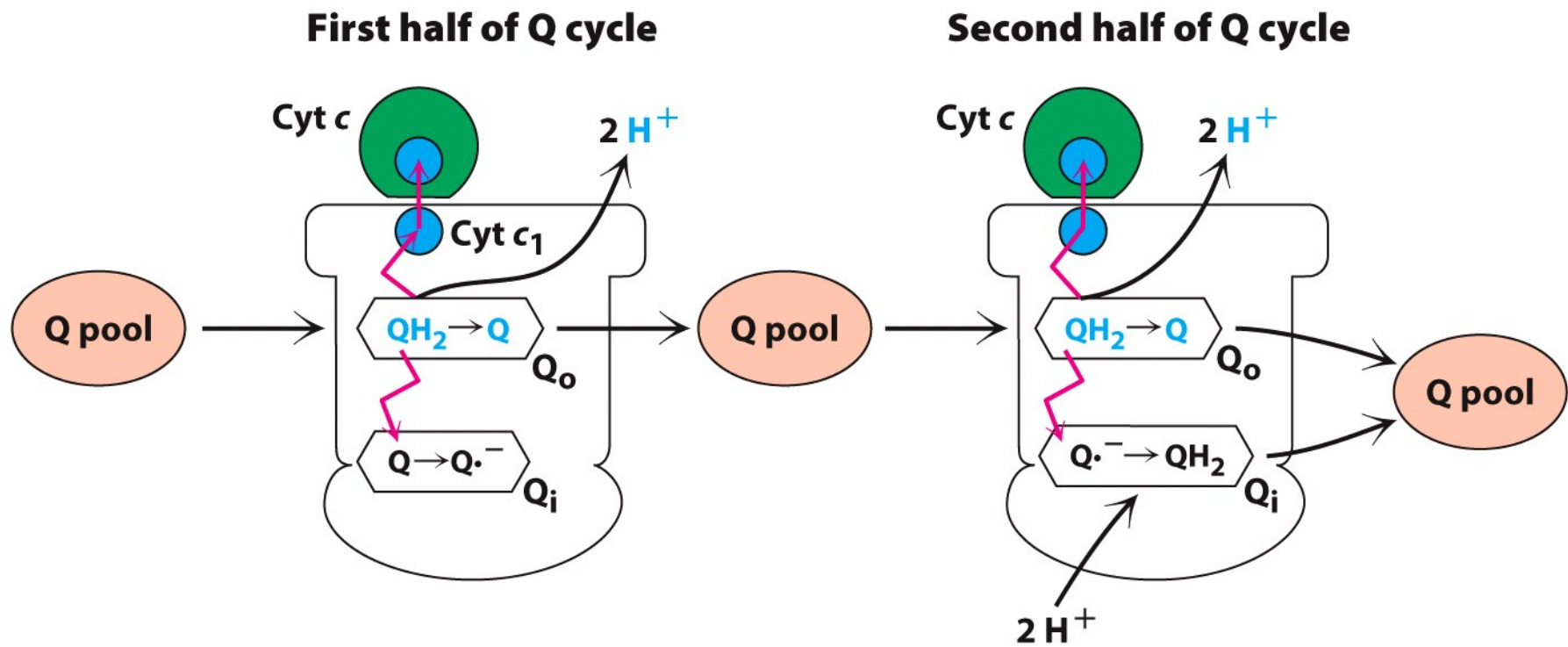


Figure 20.12
Biochemistry: A Short Course, Second Edition
 © 2013 W. H. Freeman and Company

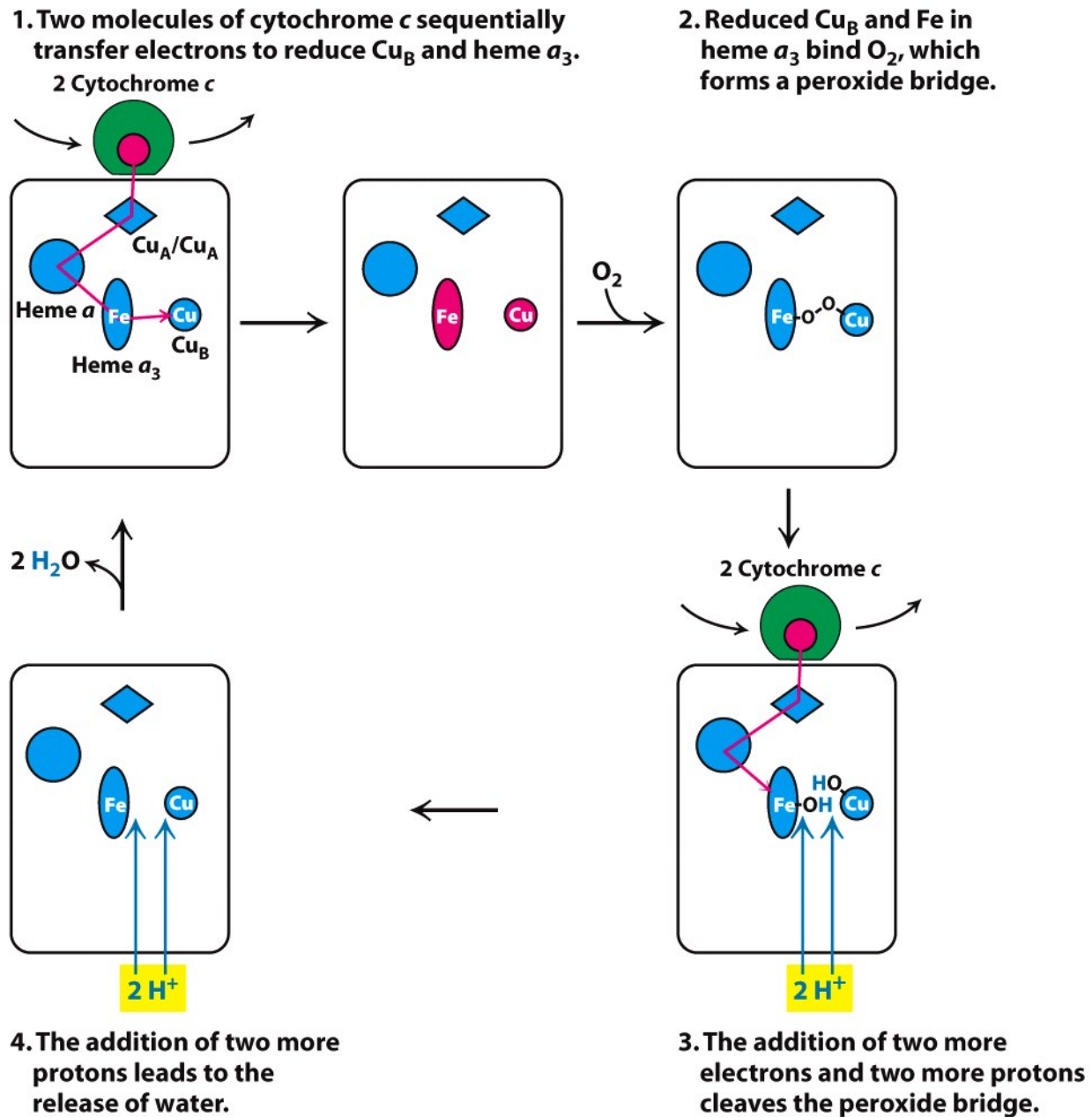


Figure 20.13
Biochemistry: A Short Course, Second Edition
© 2013 W. H. Freeman and Company

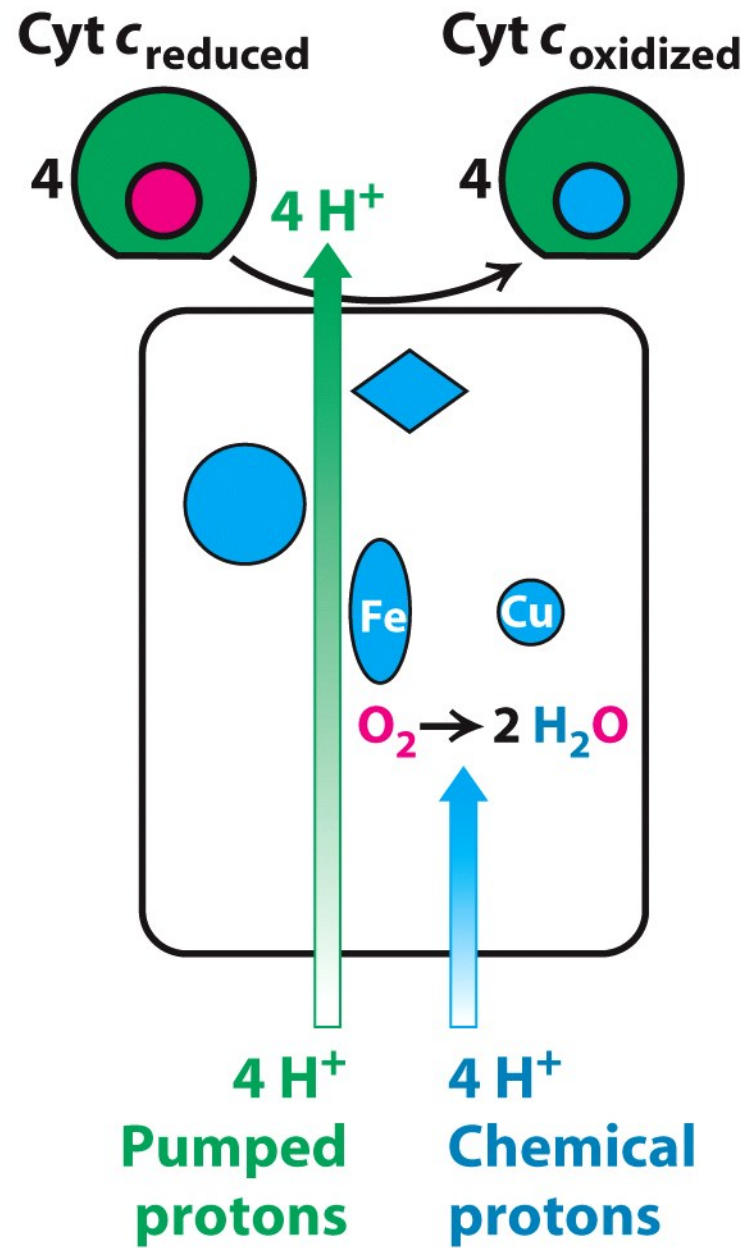


Figure 20.14
Biochemistry: A Short Course, Second Edition
 © 2013 W. H. Freeman and Company

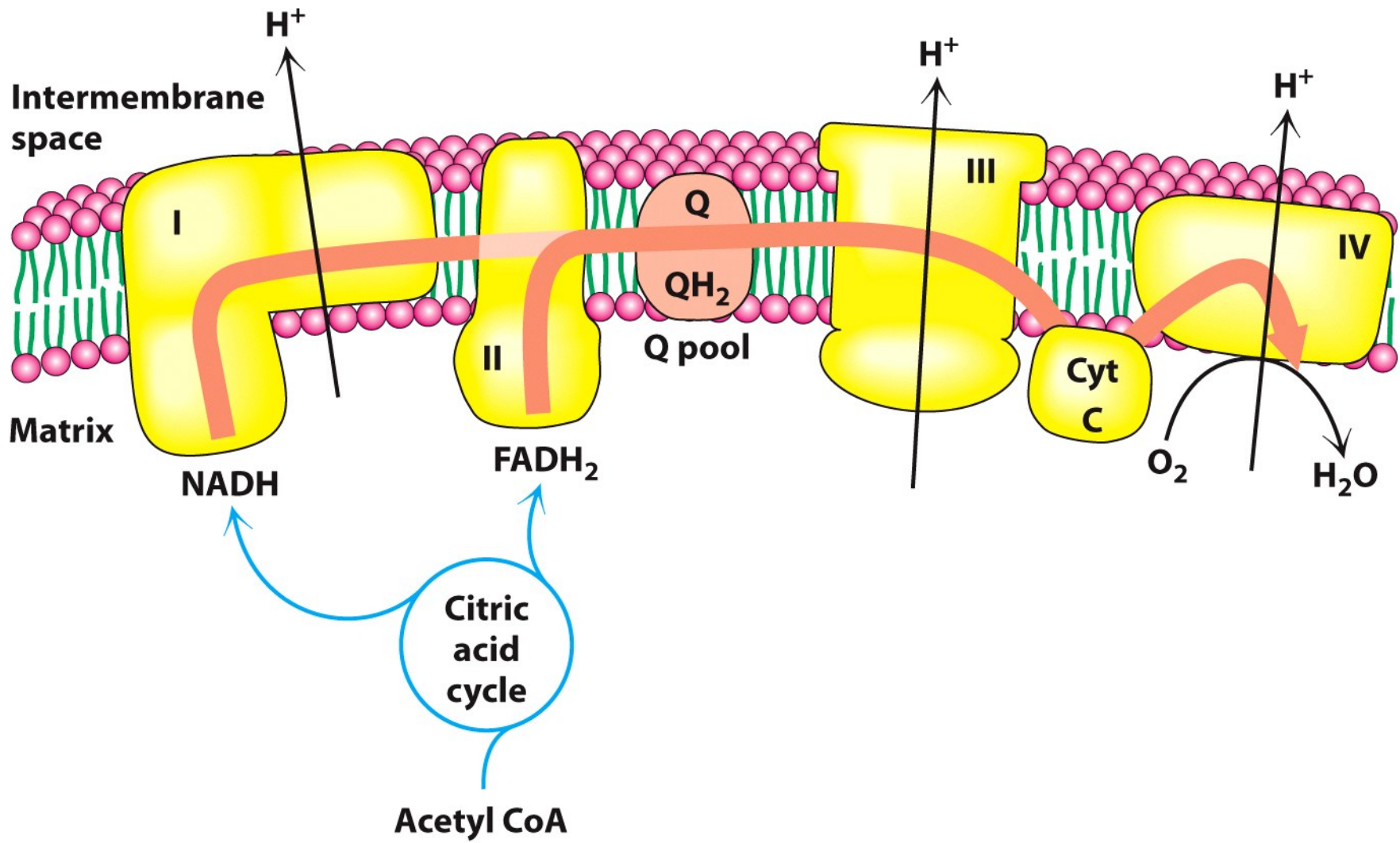


Figure 20.15
Biochemistry: A Short Course, Second Edition
 © 2013 W. H. Freeman and Company

Table 20.3 Pathological and other conditions that may be due to free-radical injury

Atherogenesis	Acute renal failure
Emphysema; bronchitis	Down syndrome
Parkinson disease	Retrolental fibroplasia (conversion of the retina into a fibrous mass in premature infants)
Duchenne muscular dystrophy	Cerebrovascular disorders
Cervical cancer	Ischemia; reperfusion injury
Alcoholic liver disease	
Diabetes	

Source: After D. B. Marks, A. D. Marks, and C. M. Smith, *Basic Medical Biochemistry: A Clinical Approach* (Williams & Wilkins, 1996), p. 331.

Table 20.3

Biochemistry: A Short Course, Second Edition

© 2013 W. H. Freeman and Company

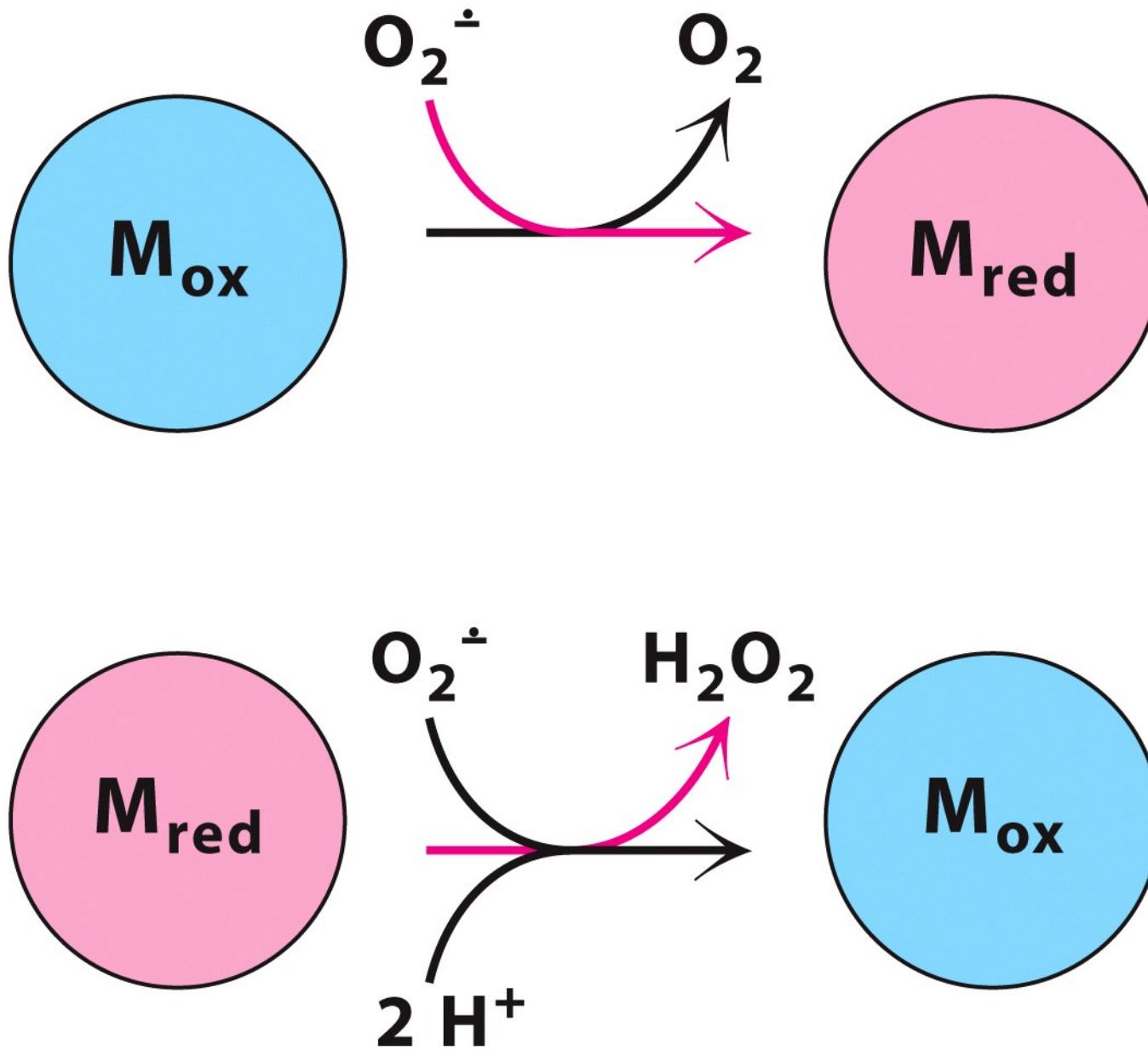


Figure 20.17
Biochemistry: A Short Course, Second Edition
© 2013 W. H. Freeman and Company